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## Limits of UV-light acceleration on the photooxidation of low-density polyethylene

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### Abstract

Lifetime prediction of polymers requires accelerated ageing based on an increase in UV light intensity and temperature. However, the representativeness of artificial ageing is still an issue. The relevance of highly accelerated UV-induced weathering raises the question of reciprocity failure depending on the polymer. The influences of irradiance and temperature on the photooxidation kinetics of polyethylene were studied to determine the extent to which increasing light intensity leads to accelerated photooxidation of polyethylene. Polyethylene films were subjected to accelerated photooxidative experiments at various UV light intensities and different temperatures, and the amounts of carbonylated photoproducts were monitored by IR spectroscopy. The results show reciprocity failure for polyethylene photooxidation and the limits of acceleration by UV light.

**Keywords:** polyethylene, photooxidation, irradiance, acceleration, reciprocity

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## 1. Introduction

The issue of lifetime prediction of polymeric materials has been a major concern for decades. Laboratory studies based on artificial ageing are required for proper determination of the long-term durability of polymers. Despite the long-term effort of various laboratories and industrial R&D centres, the reliable and accurate prediction of polymer lifetime is still an issue. Acceleration of UV light-induced degradation of polymers requires devices which allow working at intensities much higher than the solar light intensity. Experiments for predicting the environmental durability are based on accelerated artificial ageing methods, in which the acceleration is usually obtained by increasing the intensity of the light impinging on the surface and increasing the temperature of the exposed samples. Testing in these accelerated conditions can provide useful information in shorter time periods, but the stresses (light, heat,...) and their intensity have to be chosen accurately to avoid unrealistic degradation modes and failure mechanisms not observed in "true" environmental conditions [1,2]. There are many sources of misleading and erroneous results. Understanding the sources of errors is mandatory for developing relevant testing methods [3,4,5,6,7].

The issue of the representativeness of accelerated artificial ageing has been the object of many papers published in the scientific literature. Accurate prediction of the durability of polymers and data indicating that the material is capable of fulfilling its expected service requirement safely are of crucial importance.

Studies focusing on the influence of temperature on thermooxidative degradation have gained considerable attention for years [8]. Thermal degradation will not be considered in this article. It is often reported that the chemical processes governing thermal oxidation follow Arrhenius behaviour, although there is evidence that some polymers show curvature in the Arrhenius plot of oxidation rates [9]. It must also be considered that increasing the temperature of the samples during the experiments can cause heterogeneous degradation due to diffusion-limited oxidation. This issue is also well documented and has been reported in a large number of scientific papers [10, 11].

In the case of the reactions produced via light exposure, the situation is more complex. The accelerated ageing methods use commercial devices for accelerated weathering or homemade constructed devices. The acceleration of ageing is usually obtained by increasing the intensity of the UV light and by simultaneously increasing the temperature of the exposed samples. Once again, the main question is the representativeness of the accelerated ageing experiments for various reasons.

The mechanism of photooxidation can be wavelength-dependent. Most papers agree with the fact that light sources emitting short wavelengths, which are not present in sunlight, have to be rejected. The difference between the harshness of short and long wavelengths in the UV and visible domains of the solar radiation spectrum is well understood and documented in the literature [12, 13, 14]. In

addition to the influence of the light distribution on the mechanism, it must also be recalled that the rate of photooxidation can be wavelength-dependent and thus may vary with the spectral distribution of the irradiation.

Increasing the intensity of UV light may be the cause of errors when predicting the lifetime. Light sources with excessively high intensities (lasers) should be avoided, since this kind of source may produce biphotonic processes, triplet-triplet annihilation, or excessive radicals recombinations, producing results similar to those obtained when using sources with short wavelengths. This causes ageing mechanisms not observed in natural conditions. The light intensities used in conventional accelerated weathering devices are, however, well below the level which would cause a significant amount of these processes associated with pulsed lasers. Previous results published some years ago unambiguously showed that the photooxidative rate depends on the light intensity [15, 16, 17]. However, it remains a fact that the effect of irradiance, even in acceptable ranges of light intensity, is still the object of severe debates. Currently, there is a growing interest regarding this issue [18, 19,20,21,22] because of the technical, economic and environmental problems arising from the degradation of polymers.

Special attention must be paid to ensure that there is no change between the mechanisms of the accelerated tests and outdoor exposure. Nonlinear oxidative degradation would be expected if diffusion of oxygen into the sample, additive migration, or other thermally controlled reactions become the rate-limiting step. Although this is a well-documented area, the effect of irradiance, even if not producing heterogeneous oxidation and remaining in acceptable ranges of wavelengths and light intensity, is not understood well and is still the object of severe debates.

In the presence of light and air, light induces photochemical reactions, which can be described by a classical chain oxidation mechanism. The global photooxidation processes depend upon both the effective irradiance and the temperature. Photooxidation is thermally activated, with activation energies on the order of a few tens of kJ/mol. Increasing the temperature allows for accelerating the photooxidation processes, and it also allows maintaining a relevant balance between the photochemical processes and thermodegradation. This is important in the case of polymers susceptible to parallel modification provoked by light or by temperature [23,24,25].

The effect of irradiance on the rate of a photochemical process can be characterized by the reciprocity. A material obeys reciprocity if the degradation is a function of the total radiant energy and not a function of the rate at which the energy is applied. The reciprocity law can be described as:

$$I t = \text{constant, with } I = \text{irradiance and } t = \text{exposure time.}$$

The rates of photochemical processes as a function of light intensity usually follow the Schwarzschild law [26]:

$$k = A I^p$$

where  $k$  is the reaction rate,  $A$  is a proportionality constant,  $I$  is intensity (or irradiance), and  $p$  is the Schwarzschild coefficient, which is an experimentally derived number. This may also be written as:

$$I^p t = \text{constant}$$

When  $p = 1$ , Schwarzschild's law becomes the reciprocity law. For most (unstabilized, stabilized or pigmented) polymeric systems, the literature reports that the  $p$ -coefficient ranges between 0.5 and 1 [18]. However, it has to be considered that most reciprocity experiments reported were conducted with commercial, stabilized or pigmented materials and, more importantly, most of the studies on the effect of irradiance and/or temperature reported in the literature focus on properties such as colour change and yellowing, gloss loss, crystallinity, modulus, elongation at break, and mechanical properties [20,27,28]. These properties are interesting from a practical point of view, but the changes in these properties are not a direct consequence of UV light degradation and in most cases involve several processes, some of which are not light-dependent. A unique and clear relationship between the degradation of the polymers properties and the chemistry which is involved is still an issue [29]. This makes the understanding of degradation more complex and prevents researchers from proposing general concepts.

Therefore we focused on the influence of the extent of the applied stresses (light, temperature) on the most primary detectable events, which are the chemical reactions involved in the degradation of the properties.

Reciprocity failure occurs when the coefficient of proportionality changes with light intensity. This is critical in lifetime prediction because light is usually applied at a high and constant irradiance during testing, while in general, the usage conditions are lower irradiance and cycles of light exposure. As recalled above, in most cases, the light intensities used in conventional accelerated weathering devices are well below the level which would cause a significant amount of undesired processes.

In the present study, we focused on the photodegradation of polyethylene, a widely used polymer whose degradation mechanisms are well known. This paper reports a study of the influence of light intensity on the rate of photooxidation of polyethylene. Chemical modification provoked by UV light was characterized by infrared spectrometry measurements. We also present results on the photodegradation of bisphenol A polycarbonate (PC) and poly(ethyleneterephthalate) (PET). The long-term objective is to define the limits and accuracy of accelerated weathering of polyethylene and to discuss the relevance of "highly" accelerated UV weathering tools.

## **2. Experimental**

### **2.1. Materials**

Linear low-density polyethylene (LLDPE) was supplied by SABIC® (LLDPE 324CE, density = 924 kg.m<sup>-3</sup>, M<sub>w</sub> = 73,000 g.mol<sup>-1</sup>) (description in reference [30]).

PC was purchased from Aldrich (M<sub>w</sub> = 64,000 g.mol<sup>-1</sup>).

PET films were kindly supplied by Toray Film Europe (Miribel-France) in the form of thin, uniform films with thicknesses of 125 microns.

### **2.2. Preparation of polymer films**

LLDPE films with different thicknesses (from 50 to 130 microns) were obtained by compression-moulding of LLDPE pellets at 140° and 200 bar.

PC films (100 microns) were obtained by compression-moulding of PC pellets at 200°C and 200 bar.

### **2.3. Photothermal ageing**

The samples were irradiated in a MHE unit (Atlas/Ametek) at different temperatures ranging from 51 and 65°C at different irradiances (from 1.1 to 300 W.m<sup>-2</sup> between 300 and 415 nm – see below). Three films of the polymer sample per ageing time were prepared and characterized to ensure the reproducibility of the spectrometry measurements.

This SEPAP MHE was equipped with a medium-pressure mercury lamp located in the centre of the apparatus, around which the samples were placed on a rotating carousel. Wavelengths below 300 nm were filtered out. The irradiance between 300 and 415 nm was fixed at 90 (M mode) or 300 W.m<sup>-2</sup> (H mode). Irradiance and Black Standard Temperature (T<sub>BST</sub>) were on-rack controlled by a wireless XENOSENSIV light and a BST monitor. The temperature of the air in the device (T<sub>ch</sub>) was controlled by a captor placed inside the chamber. As the polymer films used in this study were transparent, the temperature on the surface of the films was considered equal to the chamber temperature.

The light intensity can be modulated by filters. Neutral density filters were used to reduce the amount of light impinging on the polymer samples. The use of neutral density filters (blocking wavelength range: 250-700 nm) allowed for carrying out experiments with irradiances below 90 W.m<sup>-2</sup>. Neutral density filters were placed in front of the polymer films in the MHE device working at 90 W.m<sup>-2</sup>, so that the irradiance received by the samples could be attenuated to 1.1 W.m<sup>-2</sup>. The filters had been obtained from MTO, and were 5 cm square. The measured transmission at 340 nm is shown in Table 1. The transparency of the filters was constant in the range of wavelengths 300-500 nm. It has also to be noted that the transparency was not modified after exposure to light.

Table 1 – Spectral characteristics of the neutral density filters and irradiance received by the films.

Transmittance (%) of the filter	Irradiance (W.m <sup>-2</sup> ) received by the films
1.2	1.1
11.8	10.6
14.1	12.7
25.5	23.0
46.1	41.5

#### **2.4. Characterization method**

##### IR spectroscopy

The thickness of the samples was low enough to allow for monitoring the chemical changes in transmission mode. This mode of detection offers the great advantage of quantitative values, in contrast to the ATR mode, which is sensitive to the surface quality, requires internal corrections and only yields semi-quantitative information. The spectra were recorded with a Nicolet 6700 FTIR spectrometer working with OMNIC software. The spectra were obtained using 32 scans and a 2 cm<sup>-1</sup> resolution.

#### **3. Results and discussion**

Infrared spectroscopy has proven to be a useful and is frequently used analytical technique for monitoring the oxidation process of polyethylene [31]. Figure 1 illustrates the changes to the LLDPE spectra, which occurred during the photooxidation, in the range of the carbonyl absorption.

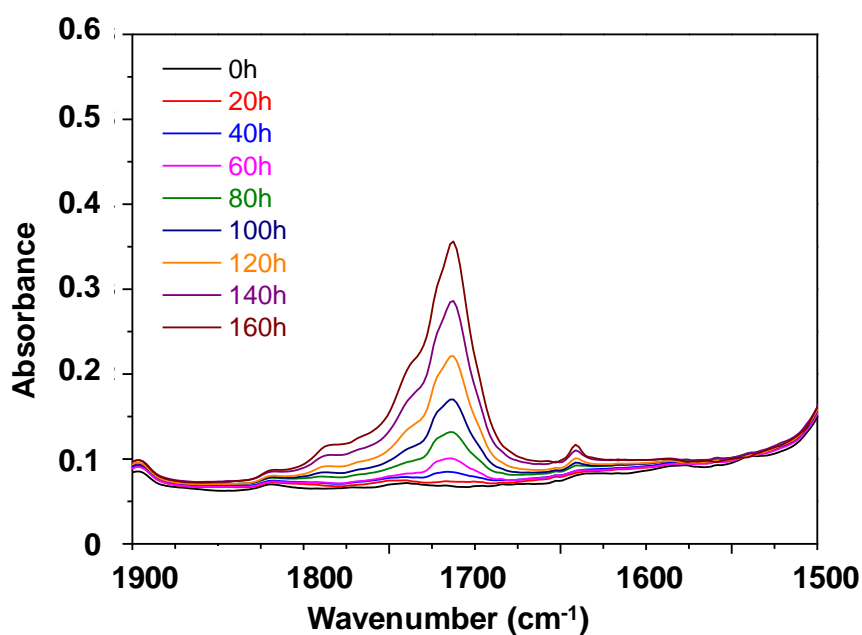


Figure 1: Infrared spectra of a LLDPE film (80 microns) at different times of photooxidation in the MHE device (90 W.m<sup>-2</sup> and 60°C)

The changes in the spectrum indicate that ketones (1718 cm<sup>-1</sup>) are formed during the initial steps, and carboxylic acids (1713 cm<sup>-1</sup>), esters (1735 cm<sup>-1</sup>) and lactones (1780 cm<sup>-1</sup>) are formed during secondary processes. One can also notice the formation of isolated double bonds (1640 cm<sup>-1</sup>). In the range of hydroxyl frequencies, not shown here, a broad band with a maximum at 3420 cm<sup>-1</sup> and a sharp absorption band at 3550 cm<sup>-1</sup> appeared in the spectrum. These bands were attributed to the formation of monomeric hydroperoxides (3550 cm<sup>-1</sup>) and hydrogen-bonded alcohols and hydroperoxides (3420 cm<sup>-1</sup>). The intensities of both bands remained very low, confirming that the stationary concentrations of hydroperoxides were rather small under the conditions of photooxidation [32]. At present, the mechanism of photooxidation accounting for the main routes of degradation of LLDPE is fairly well understood. LLDPE is a non-absorbing polymer in the sense that it does not contain chromophoric functions in its chemical structure, which can absorb in the range of solar light. The basic law of photochemistry should be recalled: light absorption is necessary to provoke photochemical reactions. In the case of non-absorbing polymers, the absorption of light is attributed to unidentified chromophores, which can vary from one LLDPE to another. Extensive work has been devoted to the nature of chromophoric species [33, 34, 35] and it is well known that oxidized species formed during processing, such as hydroperoxides or ketones, are potential



candidates. Light absorption by chromophoric defects results in the formation of radicals, which can react following different pathways: abstraction of a hydrogen atom from the macromolecular chain, addition to an unsaturated group (crosslinking reaction [36]), or addition to oxygen, which leads to the formation of various products as recalled above (Fig.1). The intensity of the band at  $1713\text{ cm}^{-1}$  is often used as a measure of the concentration of carbonyl compounds (mainly carboxylic acids) and to quantify the extent of the oxidation [35]. The slight differences in the film thicknesses between the various samples were corrected by adjusting the absorbance at  $720\text{ cm}^{-1}$ .

### 3.1. Influence of irradiance (90 and $300\text{ W.m}^{-2}$ ) at a constant temperature

We evaluated the influence of the irradiance on the oxidation of samples with a thickness of 80 microns exposed to 90 and  $300\text{ W.m}^{-2}$  in the chamber of the MHE device at a sample temperature of  $55^\circ\text{C}$ . The absorbance at  $1713\text{ cm}^{-1}$  was measured and plotted as a function of the exposure time. Figure 2 shows the obtained results.

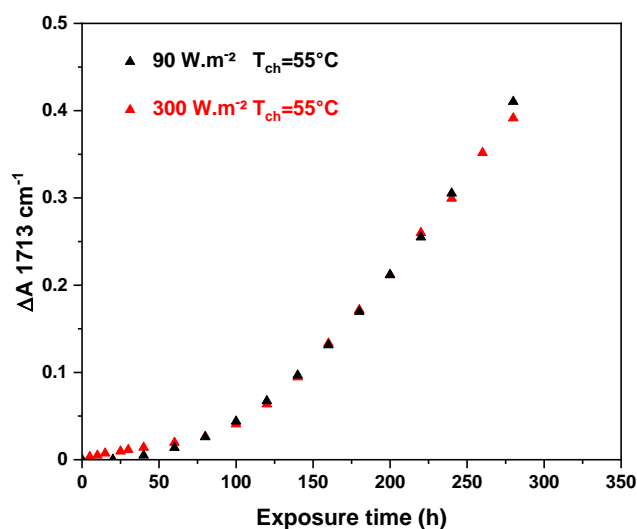


Figure 2: Carbonyl absorbance  $\Delta A$  ( $1713\text{ cm}^{-1}$ ) vs. ageing time as a function of irradiance ( $90$  and  $300\text{ W.m}^{-2}$ ) at a sample temperature of  $55^\circ\text{C}$  (LLDPE film thickness 80 microns).

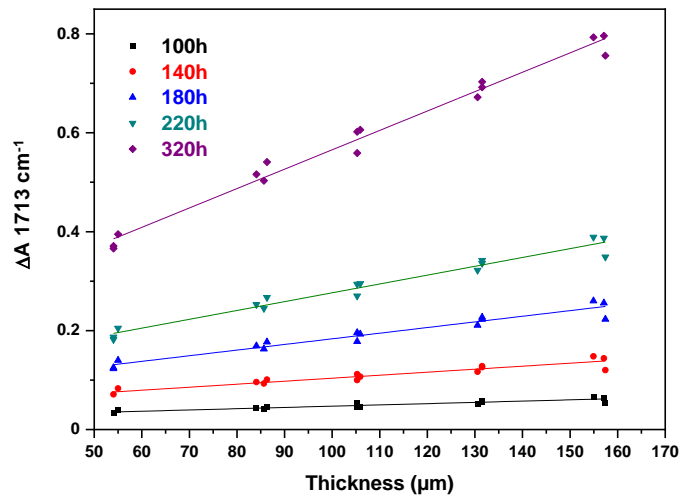
Unexpectedly, it was observed that the kinetic curves obtained from exposing the samples to two different irradiances ( $90$  and  $300\text{ W.m}^{-2}$ ) at the same temperature were the same. This result indicated that no acceleration of the photooxidation was observed, although the irradiance increased by a factor of approximately 3.

### 3.2. Influence of the film thickness

This result could be caused by heterogeneous oxidation of the LLDPE film. Heterogeneous oxidation of polymers is a very well-known phenomenon, which has been reported many times in the literature [10, 11, 37, 38]. As recalled in the introduction, the most widespread cause of heterogeneous degradation at the macroscopic level results from oxygen diffusion-limited effects. These effects can be observed in the conditions of accelerated ageing, since the rate of oxygen consumption resulting from the oxidative processes induced by light could exceed the rate of oxygen permeation. In this case, oxidation would occur in the surface layers, whereas the core would remain practically unoxidized. Infrared spectroscopy offers the great advantage of being able to monitor the effects of heterogeneous oxidation. Two simple analytic methods can be used: one is based on the use of micro-FTIR spectroscopy and consists of analysing a microtomed shaving of the photooxidized sample obtained in a plane perpendicular to the axis of irradiation [39,40]. Another method, which was used here, consists of measuring the intensity of an oxidation band as a function of the thickness of the film [41]. If the absorbance varies linearly with the thickness (for reasonably oxidized samples), the oxidation is expected to be homogeneous in the film. In the case of heterogeneous oxidation, a deviation from linearity is observed above a definite thickness.

Here, we used the second method and exposed to irradiation a series of LLDPE samples with thicknesses varying from 50 microns to 170 microns. The samples were irradiated in the MHE device at 90 and 300 W.m<sup>-2</sup> at 55°C. As an example, Figure 3 shows the changes in the absorbance at 1713 cm<sup>-1</sup> of the LLDPE films with various thicknesses during photooxidation in the MHE device irradiated at 90 W.m<sup>-2</sup> after various times of exposure.

a)



b)

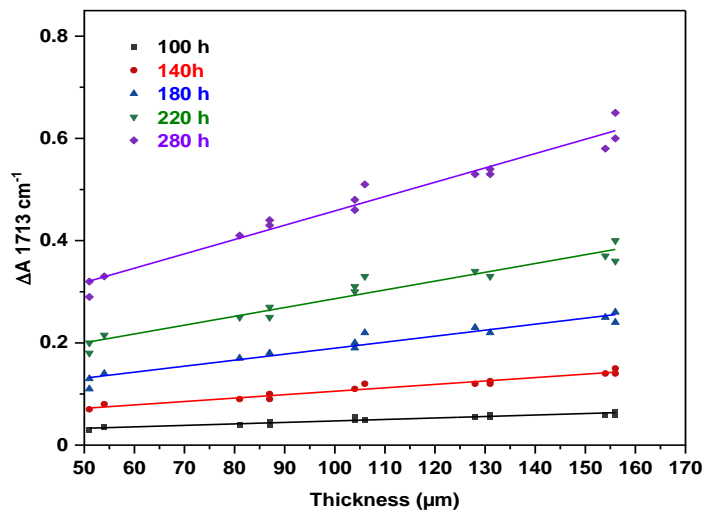


Figure 3: Changes in absorbance at  $1713\text{ cm}^{-1}$  of the LLDPE films with thicknesses from 50 to 170 microns during photooxidation in the MHE device at  $55^\circ\text{C}$  a) at  $90\text{ W.m}^{-2}$  and b) and at  $300\text{ W.m}^{-2}$

This graph shows that the absorbance at  $1713\text{ cm}^{-1}$  after various exposure times from 100 to 320 hours varied linearly with the thickness of the samples. We obtained exactly the same curves in the case of exposures at  $90\text{ W.m}^{-2}$  (Fig. 3a) and at  $300\text{ W.m}^{-2}$  (Fig. 3b). The intercepts of the curves presented in these figures are not equal to 0, which is a consequence of photo-initiating oxidation products concentrated in a thin layer of the order of 10 microns thick near the surface, as a

consequence of a surface oxidation during film fabrication, as already reported by Gugumus [42]. This result indicates that the oxidation rate was not controlled by the diffusion of oxygen into the sample.

### 3.3. Influence of the temperature (51 to 65°C)

We then examined the influence of the temperature of the chamber on the rate of photothermal degradation. The samples were exposed to either 90 or 300 W.m<sup>-2</sup> light, and the temperature was varied from 51 to 65°C. The kinetic curves of the oxidation are shown in Figure 4.

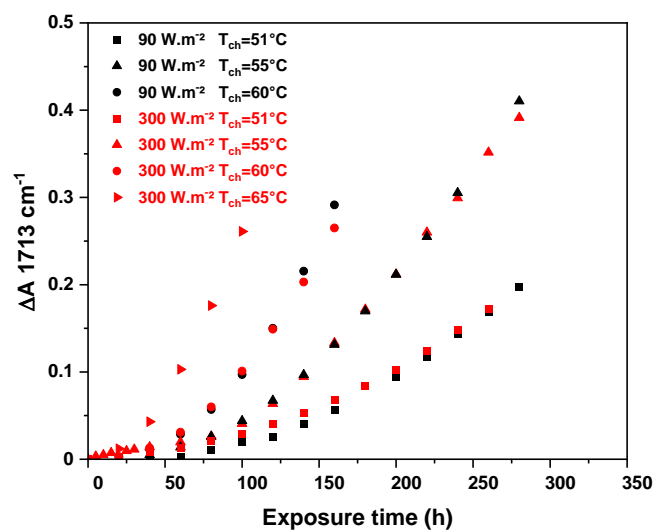


Figure 4: Absorbance  $\Delta A$  (1713 cm<sup>-1</sup>) vs. exposure time as a function of irradiance (90 and 300 W.m<sup>-2</sup>) at a sample temperature varying from 51 to 65°C (film thickness 80 microns).

The results in Figure 4 show that increasing the temperature during photooxidation experiments led to a marked increase in the rate of oxidation of the sample, which reflected the thermal activation of the photooxidation. From this graph, a shift factor can be easily calculated using the curve at the lowest temperature (51°C) as the reference. Plotting the logarithm of the shift factor as a function of the inverse of the temperature  $1/T$  allows determining the activation energy, which is proportional to the slope of the linear line fitted through the actual values (Figure 5). The calculated activation energy was approximately 74 kJ.mol<sup>-1</sup>. The kinetic curves at the various temperatures obtained in the case of irradiations at 90 and 300 W.m<sup>-2</sup> were exactly the same, this then gives the same activation energy for both the irradiances.

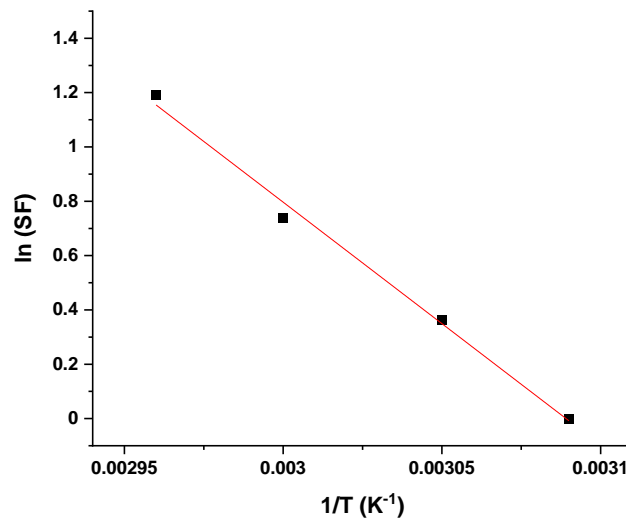


Figure 5: Arrhenius plot of the shift factor (SF) calculated from Figure 4 for LLDPE photooxidation at 300 W.m<sup>-2</sup>.

This value was different from that measured in the case of the thermooxidation of polyethylene (in the absence of UV irradiation). A previous study [30] and other studies [9, 43] indicated that, depending on the type of polyethylene (crystallinity, degree of branching, etc.) and the criteria chosen for the Arrhenius plot, the apparent activation energy  $E_a$  of the linear part of the Arrhenius curve was approximately 100-110 kJ.mol<sup>-1</sup>.

#### 3.4. Influence of irradiance (below 90 W.m<sup>-2</sup>) at a constant temperature (60°C)

The results shown in Figure 2 indicate that no increase in the photooxidation rate was observed when increasing the irradiance from 90 to 300 W.m<sup>-2</sup>, which suggests that the Schwarzschild coefficient for LLDPE is equal to 0 in the irradiance range of 90 - 300 W.m<sup>-2</sup>. The next step was to monitor the irradiance dependence of the oxidation rate at irradiances below 90 W.m<sup>-2</sup>. We performed experiments using density filters to evenly decrease the irradiance in the whole spectrum without changing the spectral distribution (see Table 1 in the experimental part). Using these filters allowed us to decrease the light irradiance well below 90 W.m<sup>-2</sup>. The obtained results are shown in Figure 6.

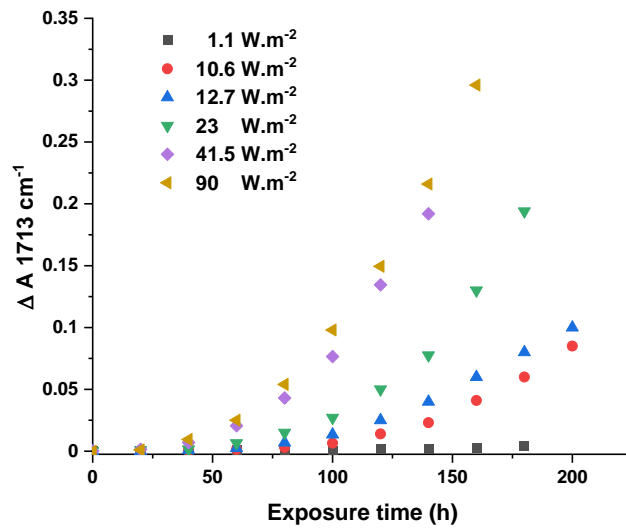


Figure 6: Variations in absorbance  $\Delta A$  ( $1713 \text{ cm}^{-1}$ ) vs. exposure time as a function of irradiance (0 to  $90 \text{ W.m}^{-2}$ ) at a sample temperature of  $60^\circ\text{C}$  (film thickness 80 microns)

These results indicated that the rate of oxidation increased with increasing irradiance, but only in the range of irradiance between  $1.1$  and  $41.5 \text{ W.m}^{-2}$ . It can be observed that the kinetics curves approach a limit when the irradiance exceeds an irradiance value of approximately  $50 \text{ W.m}^{-2}$ . This observation indicated that the Schwarzschild coefficient for LLDPE was 0 when the irradiances exceeded this value. From the kinetic curves given in Figure 6, it is possible to calculate for each irradiance a relative rate, which is defined as the time required to reach a certain absorbance at  $1713 \text{ cm}^{-1}$ , divided by the time to reach the same absorbance when irradiations were performed at  $90$  or  $300 \text{ W.m}^{-2}$  (it is recalled that the kinetic curves of oxidation were exactly similar at  $90$  or  $300 \text{ W.m}^{-2}$ ). This was done for three extents of oxidation, corresponding to absorbances  $\Delta A$  ( $1713 \text{ cm}^{-1}$ ) =  $0.05$ ,  $0.10$  and  $0.15$ . The relative rates of photothermal oxidation were then plotted against the irradiance as a log-log plot according to the equation  $k = A I^p$ . This plot is shown in Figure 7. The slope of the curves in Figure 7 was equal to  $p$ . The slope was practically the same for the three oxidation extents considered ( $\Delta A$   $1713 \text{ cm}^{-1}$  =  $0.05/0.10/0.15$ ). The slope approached  $0.5$  at irradiances below  $41.5 \text{ W.m}^{-2}$  and decreased to  $0$  when the irradiance exceeded  $90 \text{ W.m}^{-2}$ .

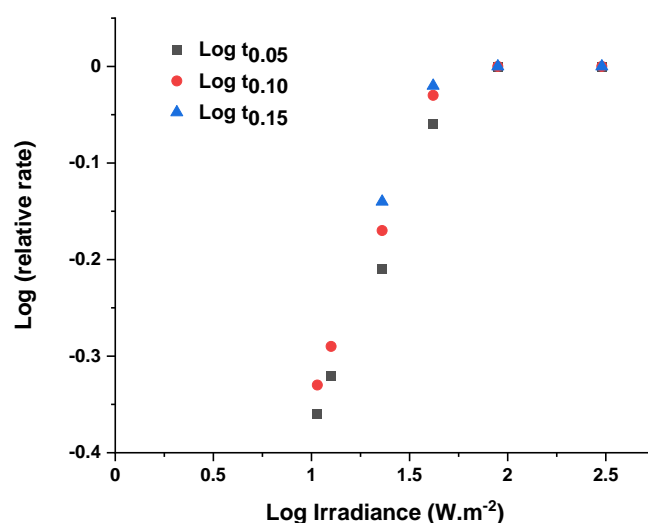


Figure 7 : Relative rates of photothermal oxidation as a log–log plot according to the equation  $k = A I^p$  (the curves correspond to  $A (1713 \text{ cm}^{-1}) = 0.05/0.10/0.15$ )

This result was unexpected, and we reproduced this experiment using two other LLDPEs of different origins. The results (not presented here) were found to be exactly similar. It is worth recalling that the results obtained in this manuscript only concern unstabilized polyethylene, and these conclusions apply only to LLDPE and do not necessarily apply to other materials with different chemical natures or materials with stabilizers.

Our results indicated that the formation of carbonyls resulting from exposure to UV light did not follow the reciprocity law. The Schwarzschild coefficient  $p$  was always below 1 and approached 0.5 at irradiances below  $41.5 \text{ W.m}^{-2}$ . More importantly, above this irradiance level, the coefficient was equal to 0, which indicated that increasing the irradiance above  $50 \text{ W.m}^{-2}$  did not provoke any increase in the oxidation rate.

A similar trend regarding the effect of light intensity has been reported in the literature [20] in the case of HDPE, which found a linear variation of the oxidation rate with the irradiance and a limit value of the oxidation rate at the highest intensities. No interpretation was given.

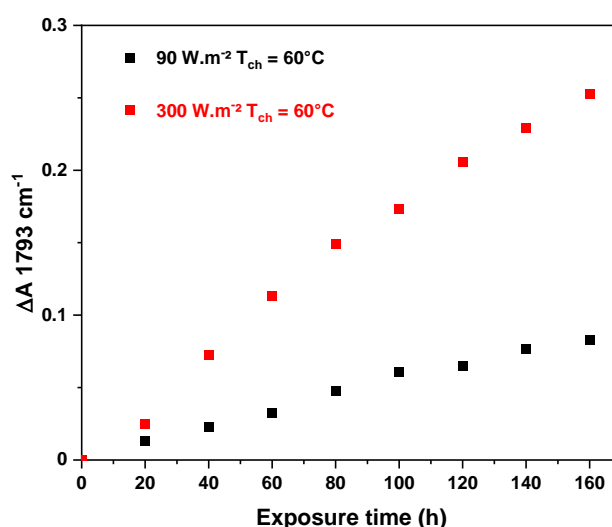
As shown above, this result does not result from a limitation of the diffusion of oxygen from the atmosphere into the samples. A plausible explanation considers that above a certain level of UV light intensity, the free radical generation is so high that recombination and/or disproportionation are faster than propagation. Similar trends have been reported in the literature by Gugumus [42] and explained by the fact that when increasing light intensity, the only reactions to be expected were photoreactions of radicals and termination reactions. This result was not expected because an

irradiance of  $90 \text{ W.m}^{-2}$  in the range of 300 and 415 nm is generally considered not too high. Moreover, the second important result is the role played by the temperature, which has not been discussed so far here. The impact of temperature on the rate of photothermal oxidation must be considered since it can fully modify the global scheme. The results shown in Figure 4 indicated that increasing the temperature increased the rate of photothermal oxidation, whereas increasing the irradiance from  $90$  to  $300 \text{ W.m}^{-2}$  had no effect on the rates. The pure photochemical reactions involved in the initiation and ramification steps are not expected to be thermally activated, whereas propagation is a thermal process. Our results suggested that increasing the temperature favoured the propagation step, after which oxidation became possible. The plausible role that the temperature could play in this process has been reported in the literature [42].

### 3.5 Influence of irradiance ( $90$ and $300 \text{ W.m}^{-2}$ ) on the photothermal oxidation of PC and PET at a constant temperature ( $60^\circ\text{C}$ )

Because the results obtained in the case of LLDPE were not expected and could be considered questionable, we repeated the experiments performed with bis-phenol A polycarbonate and poly(ethyleneterephthalate). These polymers were chosen because both have intrinsic absorptions in the UV domain of solar light, which are likely to involve direct photochemical processes. The PC and PET films were exposed to similar irradiance conditions of irradiance ( $90$  and  $300 \text{ W.m}^{-2}$ ) at a temperature of  $60^\circ\text{C}$ . The oxidation kinetic curves are shown in Figure 8a and 8b.

a)



b)



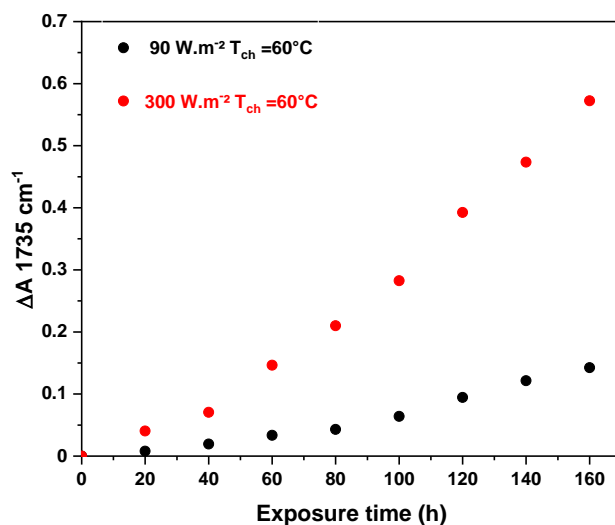


Figure 8: Variation of carbonyl absorbance vs. exposure time as a function of irradiance (90 and 300 W.m<sup>-2</sup>) at a sample temperature of 60°C (8a: PET and 8b: PC)

The p coefficients were calculated from the curves in Figure 8 using the same procedure as described above for LLDPE. In the case of PET, the p coefficient was equal to 1, which implied that the reciprocity was verified (in this range of irradiance). In the case of PC, the p coefficient was approximately 0.8, which is close to data from the literature [21, 22]. Linear increases in degradation with increased irradiance were observed for PC and poly(butylene terephthalate) irradiated in Weather-ometer at various irradiances [21] and PC was reported to follow the reciprocity law in the case of irradiations performed at various intensities in a Suntest XXL+ [22]. The comparison of polyethylene with absorbing polymers is an important result which indicates that acceleration is polymer type dependent.

#### 4. Conclusions

Accelerating degradation has been desired and explored for years, but it must be considered that the Schwarzschild coefficient p may vary from one given polymer to another, which is well known; however, it may also vary for a given polymer depending on the stability of the material brought by stabilizers. In the case of unstabilized polyethylene, no acceleration of the oxidation can be obtained by increasing the irradiance above a certain level not considered as high. Reciprocity failure with a p

from 0.5 to 0 occurs when increasing the irradiance while staying in a range of acceptable irradiance values.

On the basis of results published in the literature [42], this could be explained by the fact that recombination and/or disproportionation are faster than propagation as a result of excessive free radical generation resulting from light absorption at high irradiance levels. Acceleration of the oxidation can be obtained by increasing the temperature because propagation is thermally activated. The pure photochemical processes of polyethylene photooxidation are not thermally activated, in contrast to propagation, which does not involve light.

This result is important from a fundamental point of view, since it shows that the respective roles of light and temperature have to be discriminated. The result is also important from a practical point of view, since it provides strong indications regarding the limits of accelerated ageing and lifetime prediction. Our results indicated that there was a limit to the acceleration. Above a certain irradiance, no more acceleration occurred, but increasing the temperature allowed for an increase in the oxidation rate. However, this result could give more importance to the thermoinduced activation, which could provoke changes in the mechanism. It is indeed known that thermooxidation and photooxidation of LLDPE involve slightly different chemical mechanisms, with different consequences concerning the degradation of the functional properties of the polymer. The results obtained in the case of LLDPE highlight the issue of lifetime prediction, and a comprehensive study should include experiments performed at various irradiances and temperatures, which is time-consuming and not feasible for obvious reasons.

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## Figure Captions

**Figure 1:** Infrared spectra of a LLDPE film (80 microns) at different times of photooxidation in the MHE device (90 W.m<sup>-2</sup> and 60°C)

**Figure 2:** Carbonyl absorbance  $\Delta A$  (1713 cm<sup>-1</sup>) vs. ageing time as a function of irradiance (90 and 300 W.m<sup>-2</sup>) at a sample temperature of 55°C (LLDPE film thickness 80 microns).

**Figure 3:** Changes in absorbance at 1713 cm<sup>-1</sup> of the LLDPE films with thicknesses from 50 to 170 microns during photooxidation in the MHE device at 55°C a) at 90 W.m<sup>-2</sup> and b) and at 300 W.m<sup>-2</sup>

**Figure 4:** Absorbance  $\Delta A$  (1713 cm<sup>-1</sup>) vs. exposure time as a function of irradiance (90 and 300 W.m<sup>-2</sup>) at a sample temperature varying from 51 to 65°C (film thickness 80 microns).

**Figure 5:** Arrhenius plot of the shift factor (SF) calculated from Figure 4 for LLDPE photooxidation at 300 W.m<sup>-2</sup>.

**Figure 6:** Variations in absorbance  $\Delta A$  (1713 cm<sup>-1</sup>) vs. exposure time as a function of irradiance (0 to 90 W.m<sup>-2</sup>) at a sample temperature of 60°C (film thickness 80 microns)

**Figure 7 :** Relative rates of photothermal oxidation as a log–log plot according to the equation  $k = A I^p$  (the curves correspond to  $A$  (1713 cm<sup>-1</sup>) = 0.05/0.10/0.15

**Figure 8:** Variation of carbonyl absorbance vs. exposure time as a function of irradiance (90 and 300 W.m<sup>-2</sup>) at a sample temperature of 60°C (8a: PET and 8b: PC)

**Table 1 :** Spectral characteristics of the neutral density filters and irradiance received by the films.

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